

DETERMINATION OF THE ENTHALPY OF NEUTRALISATION

Mod. C-AV-17/EV

AC

DESCRIPTION

When a strong acid is neutralised with a strong base in a diluted solution, the same amount of heat is always released. If the reaction takes place in isobaric conditions, this heat is called enthalpy of neutralisation. The chemical reaction that generates this heat is the reaction between protons and ions and hydroxide ions to form non-ionised water. In this experiment the temperature variation is measured during the neutralisation of a solution of potassium hydroxide with a solution of hydrochloric acid and the enthalpy of neutralisation is calculated.

TRAINING PROGRAM

- Enthalpy of neutralisation
- Calorimetry
- Thermal capacity

COMPONENTS

- Calorimetric set
- Pipettes
- Pipetter
- Precision scale
- Flask
- Beakers
- Potassium hydroxide
- Hydrochloric acid
- Distilled water
- Various glassware



SUPPLIED WITH

THEORETICAL - EXPERIMENTAL HANDBOOK

