

CONDUCTIVITY OF STRONG AND WEAK ELECTROLYTES

Mod. C-AV-37/EV

DESCRIPTION

You can differentiate between strong and weak electrolytes by measuring the electrical conductivity. Strong electrolytes follow the law of Kohlrausch while weak ones are described by Ostwald's dilution law. The examination of the dependence of conductivity on the concentration lets you determine the molar conductivity of electrolytes infinitely diluted and facilitates the calculation of the degree of dissociation and the dissociation constants of weak electrolytes.

TRAINING PROGRAM

- Kohlrausch's law
- Equivalent conductivity
- Dependence of conductivity on temperature
- Ostwald's dilution law

COMPONENTI

- Magnetic stirrer
- 2 beakers, tall form, 150 ml
- 1 beaker, tall form, 250 ml
- 4 flasks, 250 ml
- 4 flasks, 500 ml
- 6 flasks, 1000 ml
- Glass funnel
- Stirrer bead
- Glass stirrer
- Pipettes, 1, 2, 25 and 100 ml
- 3 way pipette bulb
- Pasteur Pipettes
- Support base with rod
- Clamp holder
- Support clamp
- Spoon spatula
- Glass crystallizing dish
- Wash bottle
- Laboratory balance 120 g, 0.001 g
- Chemical reagents:
 - Acetic acid 1M
 - Potassium chloride
 - Distilled water



REQUIRED (NOT INCLUDED)

- EVLAB DATALOGGER mod. EVS-EXP/EV including SOFTWARE EVLAB WORKSPACE mod. SW-C-AV-37/EV for a total control of interactive experiments
- Temperature sensor mod. EVS-15/EV
- Conductivity sensor mod. EVS-BIO-07/EV
- PERSONAL COMPUTER



SUPPLIED WITH

THEORETICAL - EXPERIMENTAL HANDBOOK

